CYU p 254 #2-9, 16

2. One distinguishing property of Alkali Metals: They are VERY reactive to air and water, so not found on their own in nature.

3. Three uses of alkali metals: 1) Sodium is used in the manufacture of other chemicals (eg. NaCl is table salt), 2) Lithium in batteries 3) Potassium in fertilizers

4. Three uses of alkali earth metals: Calcium is - used to make chalk, and is in Limestone, a building material, Magnesium is used in fireworks.

5. Distinguishing property of Nobel Gases: They are INERT- very Unreactive.

6. Three applications of Nobel gases: Helium – fill balloons, Neon to make signs, Argon and Krypton in incandescent bulbs.

7. Halogens have bright colours when in gas form.

8. Industrial use of chlorine: Used in PVC to make plastics, also in bleach.

9. Hydrogen is a family of one because it can either lose or gain that one electron, allowing it to bind to almost anything.

16. Hindenburg Disaster occurred because the blimp was made for Helium, not flammable hydrogen.